

CHAPTER OVERVIEW

15: Chemical Equilibrium

In previous science classes, you may have learned that one way to distinguish chemical changes from physical changes is that physical changes—such as the melting and freezing of water—are reversible, but that chemical changes are not. In this chapter, we will see that this simple answer is not necessarily what it seems.

[15.2: The Rate of a Chemical Reaction](#)

[15.3: The Idea of Dynamic Chemical Equilibrium](#)

[15.4: The Equilibrium Constant - A Measure of How Far a Reaction Goes](#)

[15.5: Heterogeneous Equilibria- The Equilibrium Expression for Reactions Involving a Solid or a Liquid](#)

[15.6: Calculating and Using Equilibrium Constants](#)

[15.7: Disturbing a Reaction at Equilibrium- Le Châtelier's Principle](#)

[15.8: The Effect of a Concentration Change on Equilibrium](#)

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[15.10: The Effect of Temperature Changes on Equilibrium](#)

[15.11: The Solubility-Product Constant](#)

[15.12: The Path of a Reaction and the Effect of a Catalyst](#)

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